



Chemistry

Time Remaining: 45/45 (Minutes)**Q.1****Test 2 Atomic Structure****Chemistry Unit Wise**

The total number of electrons in a shell are calculated by:

a. $2(n)$ **b. $(n)^2$** **c. $2 \times X$** **d. $2(n)^2$** **STAR INSTITUTE LAHORE**[Click Here if Image Doesn't Load](#)**Correct Answer:**☐ A ☐ B ☐ C ☐ D**Next**



Time Remaining: 44/45 (Minutes)

Q.2

Test 2 Atomic Structure

Chemistry Unit Wise

If the value of $l = 3$ then the electron is located in _____ shell?

- a. K b. M
c. N d. L

STAR INSTITUTE LAHORE

[Click Here if Image Doesn't Load](#)

Correct Answer:

- ☐ A ☐ B ☐ C ☐ D

Next

Back



Time Remaining: 44/45 (Minutes)

Q:3

Test 2 Atomic Structure

Chemistry Unit Wise

After filling d orbitals the next electrons will enter into _____ orbital:

- a. s b. p
c. d d. f

STAR INSTITUTE LAHORE

[Click Here if Image Doesn't Load](#)

Correct Answer:

- ☐ A ☐ B ☐ C ☐ D

Next

Back



Time Remaining: 44/45 (Minutes)

Q.4

Test 2 Atomic Structure

Chemistry Unit Wise

An element form M^{+3} ion and belong to 3rd period of periodic table the number of protons in its nucleus are?

a. 31

b. 15

c. 13

d. 11

STAR INSTITUTE LAHORE

[Click Here if Image Doesn't Load](#)

Correct Answer:



A



B



C



D

Next

Back



Time Remaining: 44/45 (Minutes)

Q.5

Test 2 Atomic Structure

Chemistry Unit Wise

According to Aufbu's principal the highest energy orbital will be filled:

- a. Immediately b. initially
c. in the end d. first

STAR INSTITUTE LAHORE

[Click Here if Image Doesn't Load](#)

Correct Answer:

- ☐ A ☐ B ☐ C ☐ D

Next

Back



Time Remaining: 44/45 (Minutes)

Q.6

Test 2 Atomic Structure

Chemistry Unit Wise

Lowest energy electron are present in:

- a. s orbital b. p orbital
c. d orbital d. forbital

STAR INSTITUTE LAHORE

[Click Here if Image Doesn't Load](#)

Correct Answer:

- ☐ A ☐ B ☐ C ☐ D

Next

Back



Time Remaining: 44/45 (Minutes)

Q.7

Test 2 Atomic Structure

Chemistry Unit Wise

The lightest particle in nucleus is?

- a. Proton
- b. Electron
- c. Neutron
- d. All have same mass

STAR INSTITUTE LAHORE

[Click Here if Image Doesn't Load](#)

Correct Answer:

☐ A ☐ B ☐ C ☐ D

Next

Back



Time Remaining: 44/45 (Minutes)

Q.8

Test 2 Atomic Structure

Chemistry Unit Wise

All orbitals of a d-sub shell are represented with four lobes except:

a. d_{xy}

b. $d_{x^2-y^2}$

c. d_{z^2}

d. d_{xz}

STAR INSTITUTE LAHORE

[Click Here if Image Doesn't Load](#)

Correct Answer:

☐ A ☐ B ☐ C ☐ D

Next

Back



Time Remaining: 44/45 (Minutes)

Q.9

Test 2 Atomic Structure

Chemistry Unit Wise

The electrons should be filled in the order of increasing energy values is according to:

- a. Pauli Exclusion Principle
- b. Hund's rule
- c. Aufbau Principle
- d. Planck's quantum theory

STAR INSTITUTE LAHORE

[Click Here if Image Doesn't Load](#)

Correct Answer:

☐ A ☐ B ☐ C ☐ D

Next

Back



Time Remaining: 44/45 (Minutes)

Q.10

Test 2 Atomic Structure

Chemistry Unit Wise

The ionization of an atom is:

- a. Always exothermic process
- b. May or may not be endothermic
- c. Always endothermic process
- d. May be exothermic or may be endothermic process

STAR INSTITUTE LAHORE

[Click Here if Image Doesn't Load](#)

Correct Answer:

☐ A ☐ B ☐ C ☐ D

Next

Back



Time Remaining 43/45 (Minutes)



Test 2 Atomic Structure

Chemistry Unit Wise

All of the following pairs are isoelectronic except

a. S^{2-} and K^+

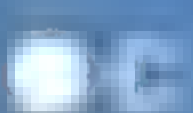
b. F^- and Ne

c. NO and N_2^-

d. C_3H_8 and CO_2

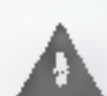
STAR INSTITUTE LAHORE

Correct Answer:



Next

Back



Time Remaining 43/45 (Minutes)



Test 2 Atomic Structure

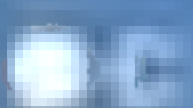
Chemistry Unit Wise

Alpha rays are actually

- a. 1 protons 2 neutrons
- b. 2 protons 2 electrons
- c. 2 protons 2 neutrons
- d. 2 protons 1 neutrons

STAR INSTITUTE LAHORE

Correct Answer:



Next

Back



Time Remaining 43/45 (Minutes)



Test 2 Atomic Structure

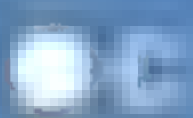
Chemistry Unit Wise

Their e/m ratio resembles with that of electrons

- a. Alpha rays
- b. Beta rays
- c. Gamma rays
- d. X-rays

STAR INSTITUTE LAHORE

Correct Answer:



Next

Back



Time Remaining 43/45 (Minutes)



Test 2 Atomic Structure

Chemistry Unit Wise

The increasing penetration effect of atomic orbitals is:

- a. $d < p < s < f$ b. $p < s < d < f$
c. $s < f < p < d$ d. $f < d < p < s$

STAR INSTITUTE LAHORE

Correct Answer:

☒ A ☐ B ☐ C ☐ D

Next

Back



Time Remaining 43/45 (Minutes)



Test 2 Atomic Structure

Chemistry Unit Wise

Which have better penetrating power?

- a. Alpha rays
- b. Beta rays
- c. Gamma rays
- d. X-rays

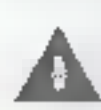
STAR INSTITUTE LAHORE

Correct Answer:

☒ A ☐ B ☐ C ☐ D

Next

Back



Time Remaining 43/45 (Minutes)

Q.16

Test 2 Atomic Structure

Chemistry Unit Wise

If $n = 3$, $l = 1$, $m = +1, 0, -1$ then orbital is:

a. 2s

b. 2p

c. 3p

d. 3d

STAR INSTITUTE LAHORE

Correct Answer:

☒ A☐ B☐ C☐ D

Next

Back

**Time Remaining 43/45 (Minutes)****Test 2 Atomic Structure****Chemistry Unit Wise**

The element shows two valency if there is sufficient gap between:

- a. Third ionization energy and fourth ionization energy
- b. First ionization energy and second ionization energy
- c. Second ionization energy and third ionization energy
- d. Fourth ionization energy and fifth ionization energy

STAR INSTITUTE LAHORE**Correct Answer:****A B C D****Next****Back**



Time Remaining 43/45 (Minutes)



Test 2 Atomic Structure

Chemistry Unit Wise

If proton number of an element 'Z' is 37 then the total number of electron in its ion ' Z^{-2} ' is:

a. 37

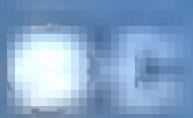
b. 35

c. 39

d. 18

STAR INSTITUTE LAHORE

Correct Answer:



Next

Back



Time Remaining 43/45 (Minutes)



Test 2 Atomic Structure

Chemistry Unit Wise

Which one of the following positive particles has maximum charge to mass ratio?

a. O^+

b. Na^+

c. K^+

d. H^+

STAR INSTITUTE LAHORE

Correct Answer:

☒ A

☐ B

☐ C

☐ D

Next

Back



Time Remaining 43/45 (Minutes)

Q20

Test 2 Atomic Structure

Chemistry Unit Wise

The charge one kilogram electron:

a. $1.602 \times 10^{-19} \text{ C}$

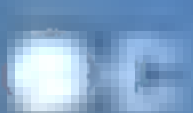
b. $1.75 \times 10^{11} \text{ C}$

c. 9.1×10^{-31}

d. 1.661×10^{-24}

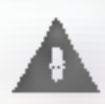
STAR INSTITUTE LAHORE

Correct Answer:



Next

Back



Time Remaining: 42/45 (Minutes)



Test 2 Atomic Structure

Chemistry Unit Wise

The relative mass of an electron is?

- a. 0 b. +1
c. 0.0005 d. -1

STAR INSTITUTE LAHORE

Correct Answer:

☒ A ☐ B ☐ C ☐ D

Next

Back



Time Remaining 42/45 (Minutes)



Test 2 Atomic Structure

Chemistry Unit Wise

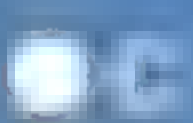
If the nucleon number for the same element is different then it refers to:

- a. difference of electron
- b. Isotopes
- c. difference of protons
- d. All of these

STAR INSTITUTE LAHORE

Correct Answer: B

Correct Answer:



Next

Back



Time Remaining 42/45 (Minutes)



Test 2 Atomic Structure

Chemistry Unit Wise

Sum of proton and neutrons in an atom is called its :

- a. isotope
- b. Atomic number
- c. Nucleon number
- a. Atomic mass

STAR INSTITUTE LAHORE

Correct Answer:

☒ A ☐ B ☐ C ☐ D

Next

Back



Time Remaining 42/45 (Minutes)

Test 2

Test 2 Atomic Structure

Chemistry Unit Wise

Electronic configuration of K is:

a. $[\text{Ar}]4s^2$

b. $[\text{Ar}]4s^1$

c. $[\text{Kr}]5s^1$

d. $[\text{He}]2s^1$

STAR INSTITUTE LAHORE

Correct Answer:

☒ A

☐ B

☐ C

☐ D

Next

Back



Time Remaining 42/45 (Minutes)



Test 2 Atomic Structure

Chemistry Unit Wise

A set of orbitals having same value of 'l' is called:

- a. Shell
- b. Sub-shell
- c. molecular orbital
- d. Energy level

STAR INSTITUTE LAHORE

Correct Answer:

- ☒ A
- ☐ B
- ☐ C
- ☐ D

Next

Back



Time Remaining 42/45 (Minutes)

Q26

Test 2 Atomic Structure

Chemistry Unit Wise

Which one of the following rule is used to arrange the sub energy levels in increasing order of energy?

- a. Hund's rule
- b. $(n+l)$ rule
- c. Octet rule
- d. Auf bau principle

STAR INSTITUTE LAHORE

Correct Answer:

☒ A ☐ B ☐ C ☐ D

Next

Back



Time Remaining 42/45 (Minutes)

open

Test 2 Atomic Structure

Chemistry Unit Wise

Which one is the heavier particle?

a. Electron

b. Proton

c. Neutron

d. Photon

STAR INSTITUTE LAHORE

Correct Answer:



Next

Back



Time Remaining 42/45 (Minutes)

Q26

Test 2 Atomic Structure

Chemistry Unit Wise

The total relative charge of an element is equal to:

- a. Its charge of electrons
- b. Zero
- c. its Charge of protons
- d. None of these

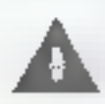
STAR INSTITUTE LAHORE

Correct Answer:

☒ A ☐ B ☐ C ☐ D

Next

Back



Time Remaining 42/45 (Minutes)



Test 2 Atomic Structure

Chemistry Unit Wise

The mass of electron is

- a. 1.6022×10^{-19} kg
- b. 1.6022×10^{-17} kg
- c. 9.1090×10^{-31} kg
- d. None

STAR INSTITUTE LAHORE

Correct Answer:



Next

Back



Time Remaining 41/45 (Minutes)

0:30

Test 2 Atomic Structure

Chemistry Unit Wise

Which one of the following determines the position of an element in the Periodic Table?

- a. chemical reactivity
- b. first ionization energy
- c. number of electrons in outer orbital
- d. number of protons in the nucleus of its atom

STAR INSTITUTE LAHORE

Correct Answer:

☒ A ☐ B ☐ C ☐ D

Next

Back



Time Remaining 41/45 (Minutes)



Test 2 Atomic Structure

Chemistry Unit Wise

An element with $4p^4$ valence electronic configuration will have period and group no. in modern periodic table?

a. 4 and IV

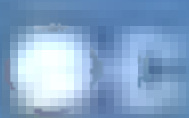
b. 4 and III

c. 4 and VI

d. 4 and V

STAR INSTITUTE LAHORE

Correct Answer:



Next

Back

**Time Remaining 41/45 (Minutes)**

Test 2 Atomic Structure

Chemistry Unit Wise

Which one of the following are Isosteres?

- a. H^{-1} and H
- b. N_2 and CO
- c. ${}_6\text{C}^{12}$ and ${}_8\text{O}^{16}$
- d. ${}_{18}\text{Ar}^{20}$ and ${}_{20}\text{Ca}^{40}$

STAR INSTITUTE LAHORE

Correct Answer:

☒ A ☐ B ☐ C ☐ D

Next

Back



Time Remaining 41/45 (Minutes)



Test 2 Atomic Structure

Chemistry Unit Wise

Which orbital is bigger in size and have maximum energy?

a. 2px

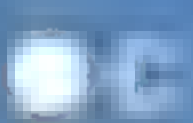
b. 3px

c. 4px

d. 5px

STAR INSTITUTE LAHORE

Correct Answer:



Next

Back



Time Remaining: 41/45 (Minutes)



Test 2 Atomic Structure

Chemistry Unit Wise

The ionic species having more electrons than neutrons is

a. Na^+

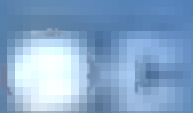
b. Mg^{+2}

c. O^{-2}

d. F^{-1}

STAR INSTITUTE LAHORE

Correct Answer:



Next

Back



Time Remaining 41/45 (Minutes)

035

Test 2 Atomic Structure

Chemistry Unit Wise

What kind of orbital must an electron with the principal quantum number $n=2$ occupy?

- a. a spherically-shaped orbital
- b. either an s or p orbital
- c. the orbital closest to the nucleus
- d. a dumb-bell-shaped orbital

STAR INSTITUTE LAHORE

Correct Answer:

☒ A ☐ B ☐ C ☐ D

Next

Back



Time Remaining 41/45 (Minutes)

Q-36

Test 2 Atomic Structure

Chemistry Unit Wise

Which property is the same for the two nuclides $\frac{40}{18}\text{Ar}$ and $\frac{40}{19}\text{K}$?

- a. the number of electrons
- b. the number of neutrons
- c. the number of nucleons
- d. the number of protons

STAR INSTITUTE LAHORE

Correct Answer:

☒ A ☐ B ☐ C ☐ D

Next

Back



Time Remaining: 41/45 (Minutes)



Test 2 Atomic Structure

Chemistry Unit Wise

A spinning electron creates

- a. magnetic field
- b. electric field
- c. quantum field
- d. none

STAR INSTITUTE LAHORE

Correct Answer:

☒ A ☐ B ☐ C ☐ D

Next

Back



Time Remaining 41/45 (Minutes)



Test 2 Atomic Structure

Chemistry Unit Wise

Atomic number of an element is 17. The number of pairs of paired and also unpaired electrons in the valence shell of atom is :

- a. 1, 3 b. 3, 1
c. 2, 2 d. 4, 1

STAR INSTITUTE LAHORE

Correct Answer:

☒ A ☐ B ☐ C ☐ D

Next

Back

**Time Remaining 40/45 (Minutes)**

Test 2 Atomic Structure

Chemistry Unit Wise

The correct set of quantum number for unpaired electron in sodium atom is:

- | | n | l | m |
|----|-----|-----|-----|
| a. | 2 | 0 | 0 |
| b. | 3 | 0 | 0 |
| c. | 2 | 1 | 1 |
| d. | 3 | 0 | 1 |

STAR INSTITUTE LAHORE

Correct Answer:

☒ A ☐ B ☐ C ☐ D

Next

Back



Time Remaining 40/45 (Minutes)

Q40

Test 2 Atomic Structure

Chemistry Unit Wise

The divisibility of atom was shown by

a. Stoney

b. J.J Thomson

c. Millikan

d. Rutherford

STAR INSTITUTE LAHORE

Correct Answer:

☒ A

☐ B

☐ C

☐ D

Submit Quiz

Back

Q. 1

The total number of electrons in a shell are calculated by:

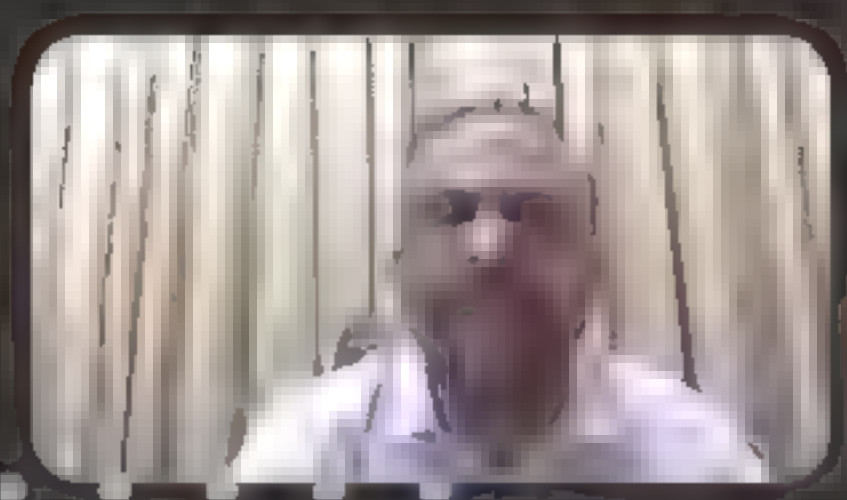
a. $2(n)$

b. $2 \times n$

c. $(n)^2$

d. $2(n)^2$

THE STAR INSTITUTE



سید اختر عباس جعفری screen

Q. 2

If the value of $l = 3$ then the electron is located in _____ shell?

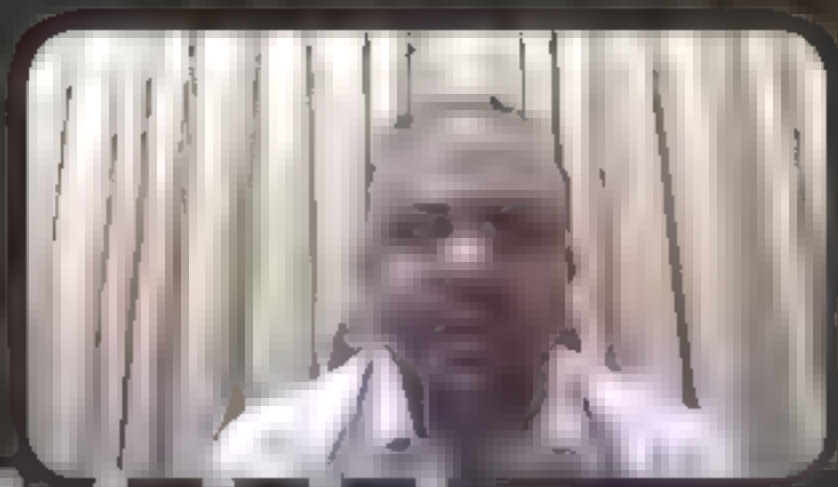
a. K

b. M

c. N

d. L

THE STAR INSTITUTE



سید اختر عباس جعفری screen

Q. 3

After filling d orbitals the next electrons will enter into _____ orbital:

a. s

b. p

c. d

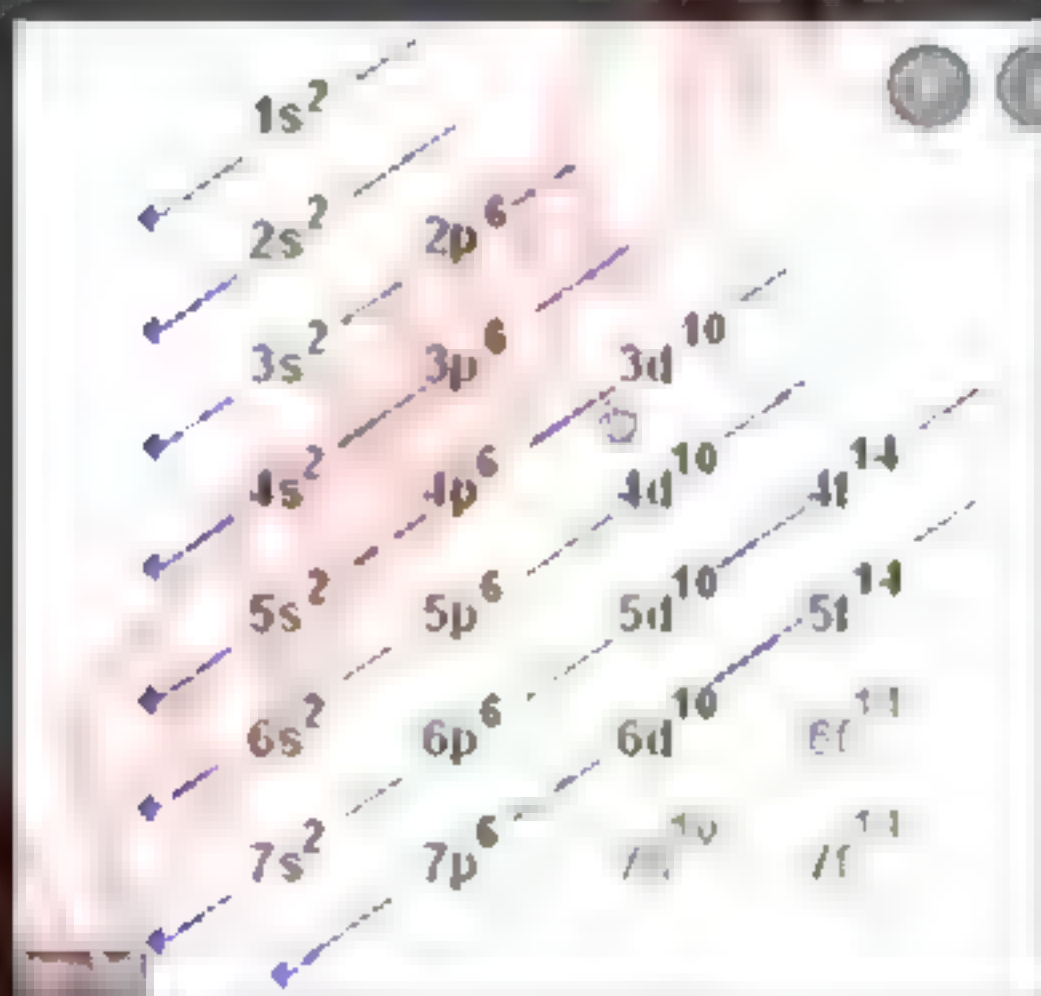
d. f

THE STAR INSTITUTE

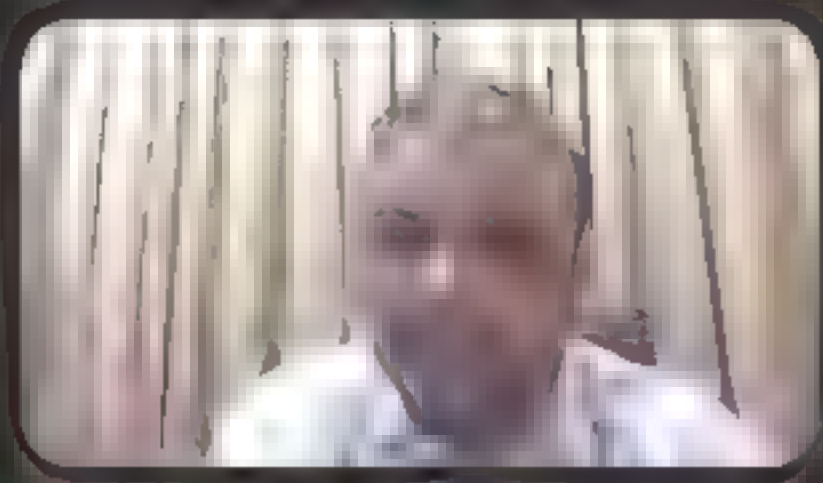


Zoom

Leave



Electron Configurations



Unmute

Start Video

Share

Participants 34

More

Q. 4

An element form M^{+3} ion and belong to 3rd period of periodic table the number of protons in its nucleus are?

a. 31

b. 15

c. 13

d. 11

THE STAR INSTITUTE

Q. 5

According to Aufbau's principal the highest energy orbital will be filled:

- a. Immediately
- b. initially
- c. in the end
- d. first

THE STAR INSTITUTE



Q. 6

Lowest energy electron are present in:

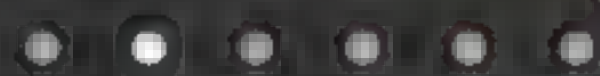
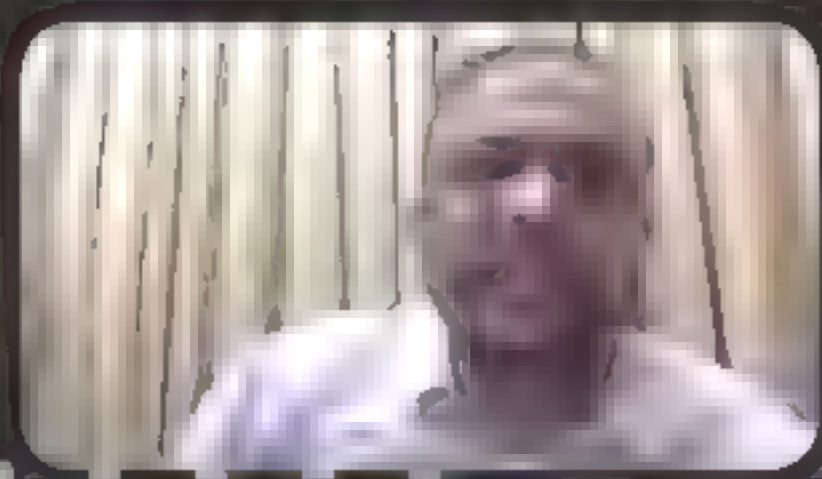
a. s orbital

b. p orbital

c. d orbital

d. f orbital

THE STAR INSTITUTE

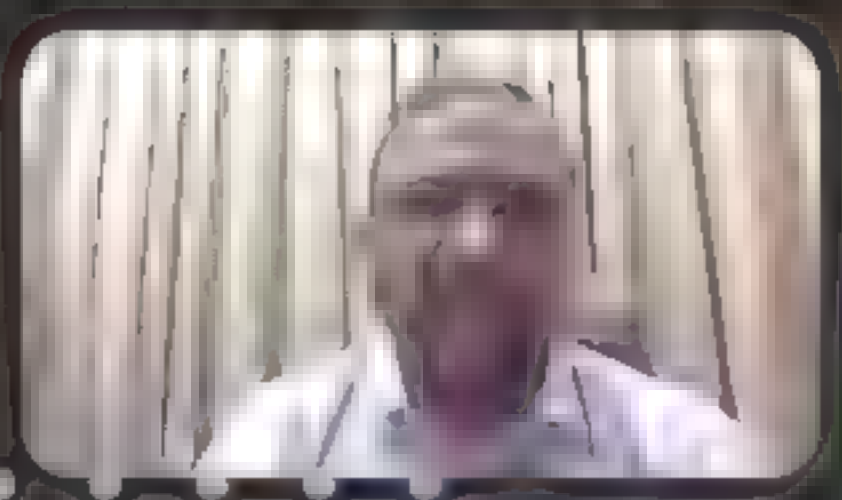


Q. 7

The lightest particle in nucleus is?

- a. Proton
- b. Electron
- c. Neutron
- d. All have same mass

THE STAR INSTITUTE





Zoom

Leave

Q. 8

All orbitals of a d-sub shell are represented with four lobes except:

a. d_{xy} b. $d_{x^2-y^2}$ c. d_{z^2} d. d_{xz}

THE STAR INSTITUTE



Unmute



Start Video



Share



Participants



More

Q. 9

The electrons should be filled in the order of increasing energy values is according to:

- a. Pauli Exclusion Principle
- b. Hund's rule
- c. **Aufbau Principle**
- d. Planck's quantum theory

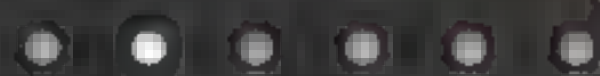
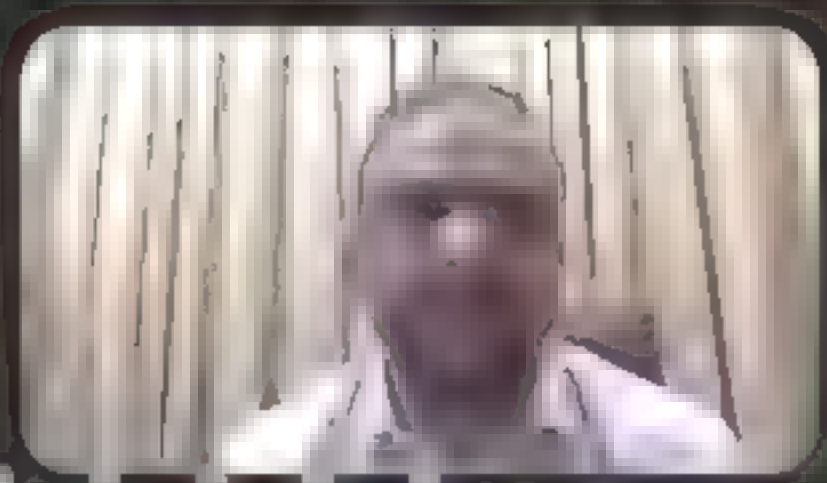
THE STAR INSTITUTE

Q. 10

The ionization of an atom is:

- a. Always exothermic process
- b. May or may not be endothermic
- c. Always endothermic process
- d. May be exothermic or may be endothermic process**

THE STAR INSTITUTE



Q. 11

All of the following pairs are isoelectronic except

a. S^{2-} and K^+

16+2 19-1

b. F and Ne

9+1 10

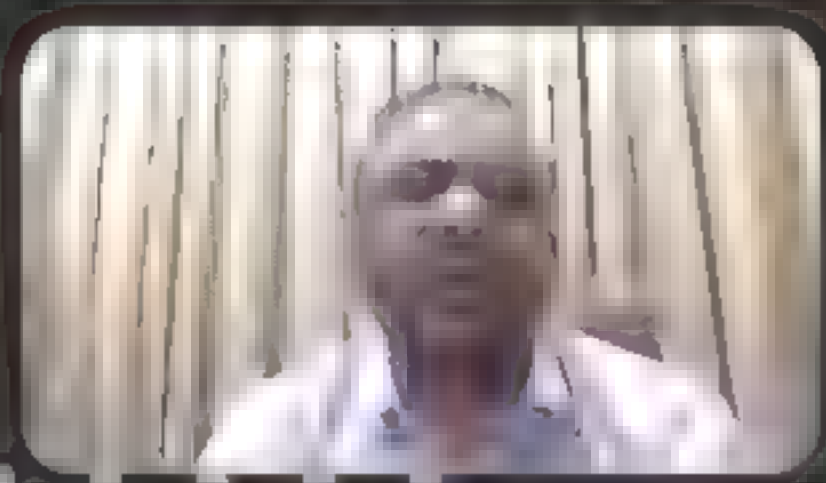
c. NO and N_2^-

7+8 7+7+1

d. C_2H_2 and CO_2

18+8 6+16

THE STAR INSTITUTE

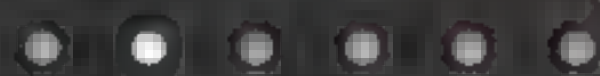
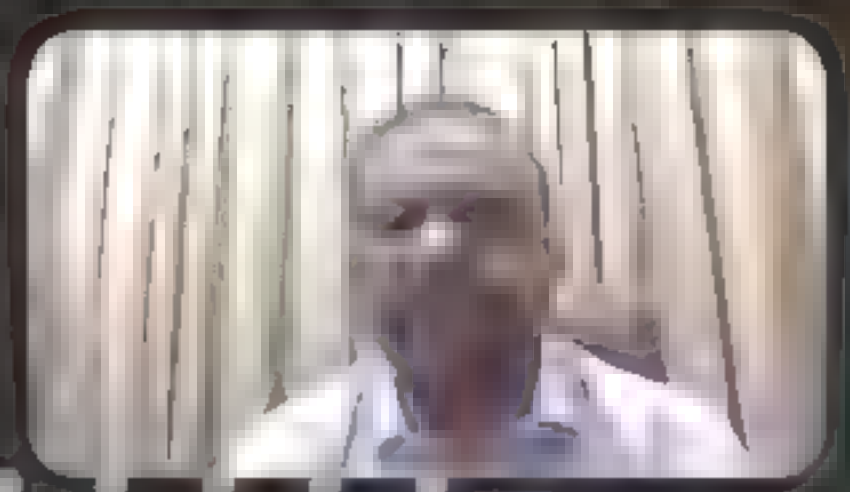


Q. 12

Alpha rays are actually

- a. 1 protons 2 neutrons
- b. 2 protons 2 electrons
- c. 2 protons 2 neutrons
- d. 2 protons 1 neutrons

THE STAR INSTITUTE



Q. 13

Their e/m ratio resembles with that of electrons

- a. Alpha rays
- b. Beta rays
- c. Gamma rays
- d. X-rays

THE STAR INSTITUTE

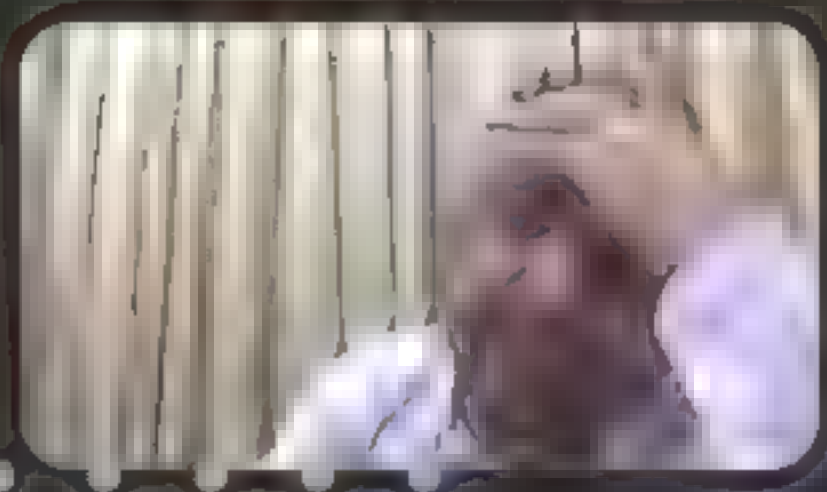


Q. 14

The increasing penetration effect of atomic orbitals is:

- a. $d < p < s < f$ b. $p < s < d < f$
c. $s < f < p < d$ d. $f < d < p < s$

THE STAR INSTITUTE

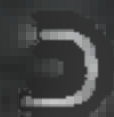
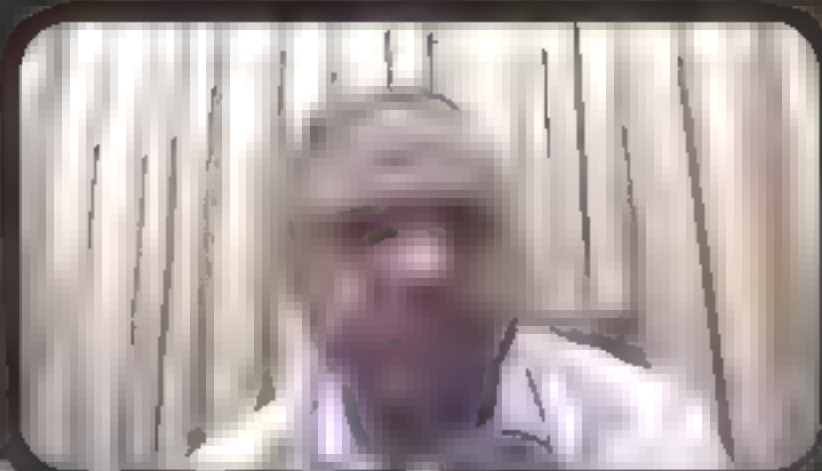


Q. 15

Which have better penetrating power?

- a. Alpha rays
- b. Beta rays
- c. Gamma rays
- d. X-rays

THE STAR INSTITUTE



Q. 16

If $n = 3$, $l = 1$, $m = +1, 0, -1$ then orbital is:

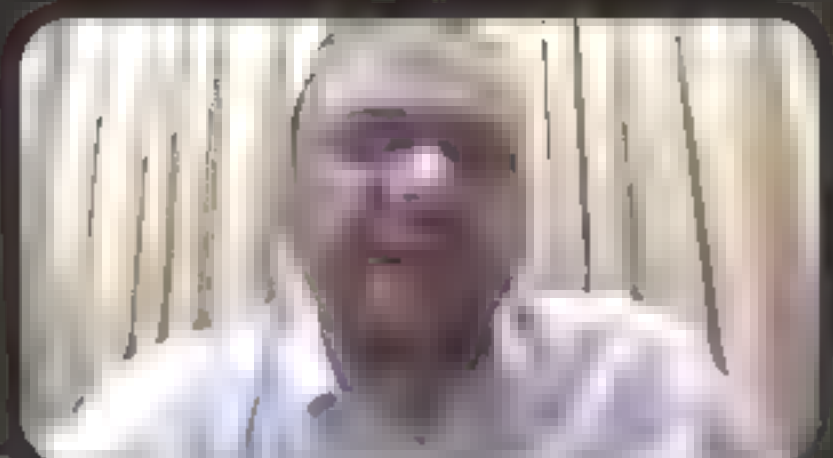
a. 2s

b. 2p

c. 3p

d. 3d

THE STAR INSTITUTE



سید اختر عباس جعفری screen

Q. 17

The element shows two valency if there is sufficient gap between:

- a. Third ionization energy and fourth ionization energy
- b. First ionization energy and second ionization energy
- c. **Second ionization energy and third ionization energy**
- d. Fourth ionization energy and fifth ionization energy

THE STAR INSTITUTE

Q. 18

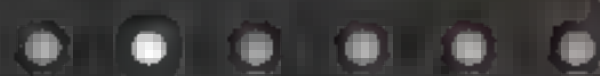
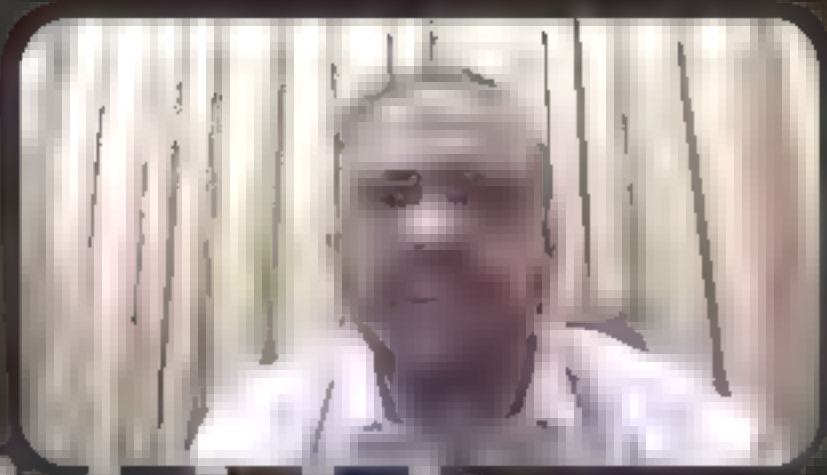
If proton number of an element Z is 37 then the total number of electron in its ion Z^{2+} is:

a. 37

b. 35

c. 39

d. 18

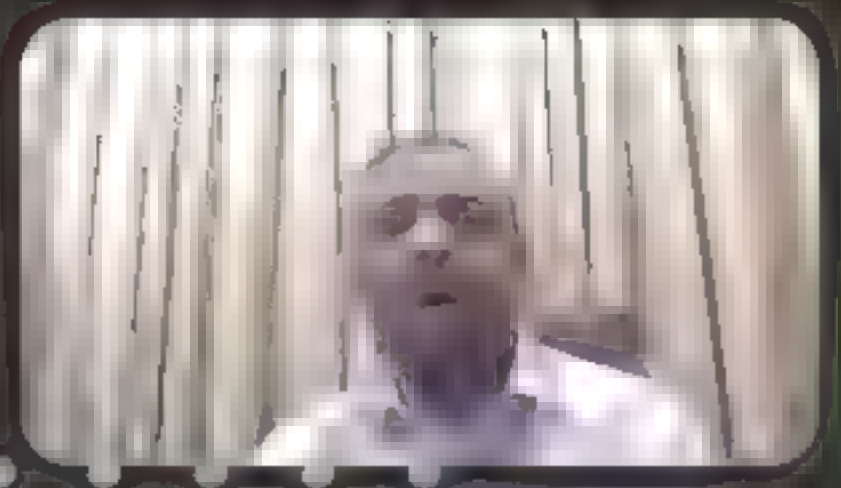
THE STAR INSTITUTE

Q. 19

Which one of the following positive particles has maximum charge to mass ratio?

a. O^{+} b. Na^{+} c. K^{+} d. H^{+}

THE STAR INSTITUTE



سید اختر عباس جعفری screen s

Q. 20

The charge one kilogram electrons:

- a. $1.602 \times 10^{-19} \text{ C}$ b. $1.75 \times 10^{11} \text{ C}$
c. 9.1×10^{-31} d. 1.661×10^{-24}

THE STAR INSTITUTE

Q. 21

The relative mass of an electron is?

a. 0

b. +1

c. 0.0005

d. -1

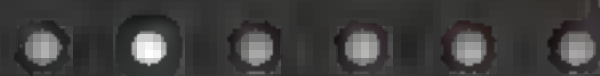
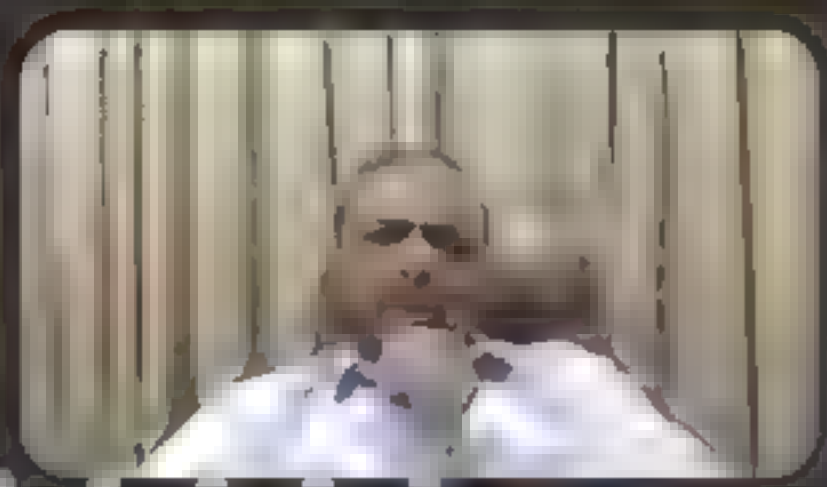
THE STAR INSTITUTE



Q. 22

If the nucleon number for the same element is different then it refers to:

- a. difference of electron
- b. Isotopes**
- c. difference of protons
- d. All of these

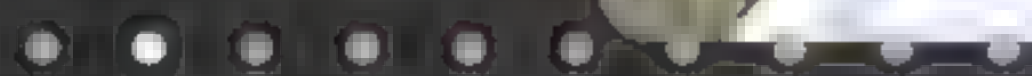
THE STAR INSTITUTE

Q. 23

Sum of proton and neutrons in an atom is called its:

- a. isotope
- ☒ b. Atomic number
- c. Nucleon number
- d. Atomic mass

THE STAR INSTITUTE



Q. 24

Electronic configuration of K is:

a. $[\text{Ar}]4s$

b. $[\text{Ar}]4s^2$

c. $[\text{Kr}]5s$

d. $[\text{He}]2s$

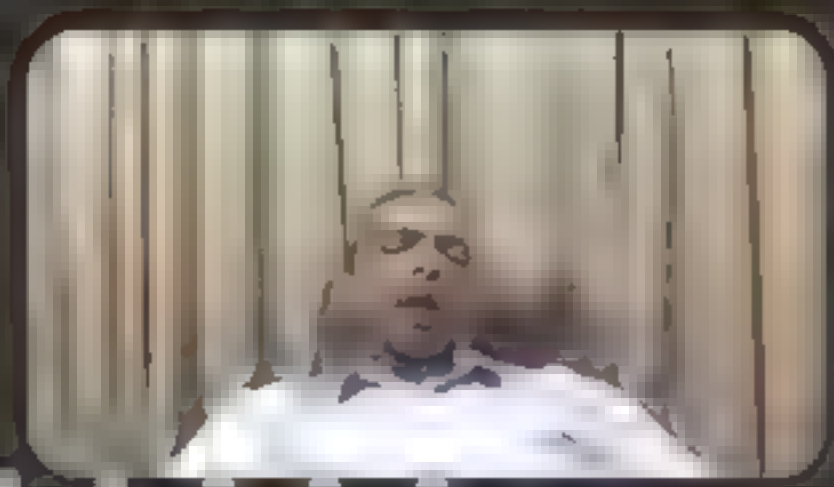
THE STAR INSTITUTE

Q. 25

A set of orbitals having same value of 'l' is called:

- a. Shell
- b. Sub-shell
- c. molecular orbital
- d. Energy level

THE STAR INSTITUTE



Q. 26

Which one of the following rule is used to arrange the sub energy levels in increasing order of energy?

a. Hund's rule

b. $(n+l)$ rule

c. Octet rule

d. Auf bau principle

THE STAR INSTITUTE

Q. 27

Which one is the heavier particle?

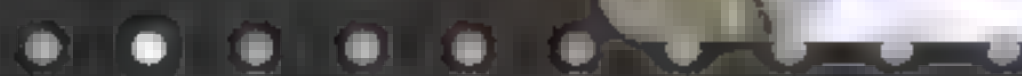
a. Electron

b. Proton

☒ c. Neutron

d. Photon

THE STAR INSTITUTE

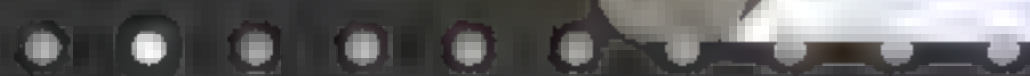


Q. 28

The total relative charge of an element is equal to:

- a. Its charge of electrons
- b. Zero
- c. Its Charge of protons
- d. None of these

THE STAR INSTITUTE



Q. 29

The mass of electron is

- a. 1.6022×10^{-19} kg
- b. 1.6022×10^{-17} kg
- c. 9.1090×10^{-31} kg
- d. None

THE STAR INSTITUTE

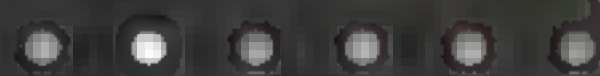


Q. 30

Which one of the following determines the position of an element in the Periodic Table?

- a. chemical reactivity
- b. first ionization energy
- c. number of electrons in outer orbital
- d. number of protons in the nucleus of its atom

THE STAR INSTITUTE



Q. 31

An element with $4p^4$ valence electronic configuration will have period and group no. in modern periodic table?

- a. 4 and IV b. 4 and III
c. 4 and VI d. 4 and V

THE STAR INSTITUTE



Q. 32

Which one of the following are Isosteres?

a. H_2 and H

b. N_2 and CO

c. C^{12} and O^{16}

d. $_{18}Ar^{40}$ and $_{20}Ca^{40}$

Isosteres: one of two or more substances (as carbon monoxide and molecular nitrogen) that exhibit similarity of some properties as a result of having the same number of total or valence electrons in the same arrangement and that consist of different atoms and not necessarily the same number of atoms.

THE STAR INSTITUTE

Q. 33

Which orbital is bigger in size and have maximum energy?

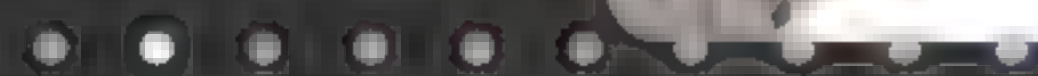
a. 2px

b. 3px

c. 4px

d. 5px

THE STAR INSTITUTE



Q. 34

The ionic species having more electrons than neutrons is

a. Na^+ b. Mg^{2+} c. O^{2-} d. F^-

THE STAR INSTITUTE



Q. 35

What kind of orbital must an electron with the principal quantum number $n=2$ occupy?

- a. a spherically -shaped orbital
- b. either an s or p orbital**
- c. the orbital closest to the nucleus
- d. a dumb-bell-shaped orbital

THE STAR INSTITUTE

Q. 36

Which property is the same for the two nuclides ${}^{40}_{18}\text{Ar}$ and ${}^{40}_{19}\text{K}$?

- a. the number of electrons
- b. the number of neutrons
- c. the number of nucleons
- d. the number of protons

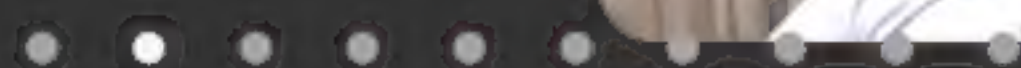
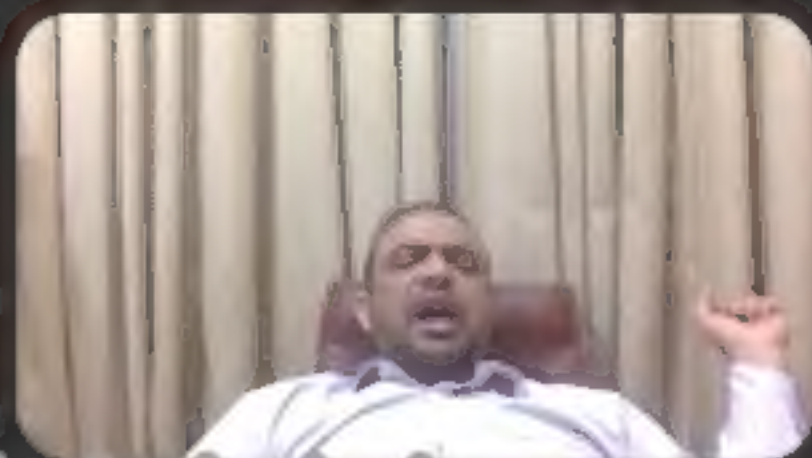
THE STAR INSTITUTE

Q. 37

A spinning electron creates

- a. magnetic field
- b. electric field
- c. quantum field
- d. none

THE STAR INSTITUTE



Q. 38

Atomic number of an element is 17. The number of pairs of paired and also unpaired electrons in the valence shell of atom is :

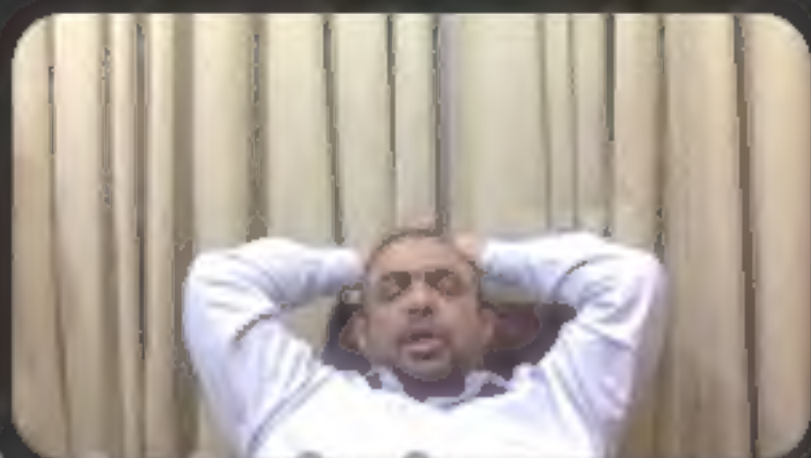
a. 1, 3

b. 3, 1

c. 2, 2

d. 4, 1

THE STAR INSTITUTE

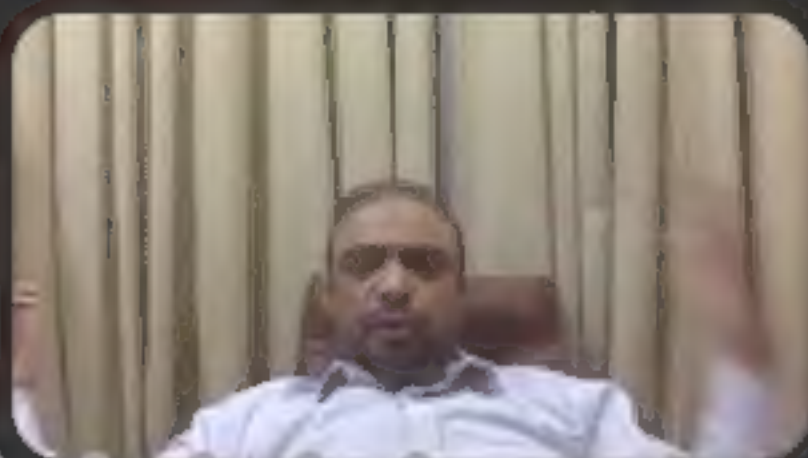


Q. 39

The correct set of quantum number for unpaired electron in sodium atom is:

- | | n | l | m |
|----|-----|-----|-----|
| a. | 2 | 0 | 0 |
| b. | 3 | 0 | 0 |
| c. | 2 | 1 | 1 |
| d. | 3 | 0 | 1 |

THE STAR INSTITUTE



Q. 40

The divisibility of atom was shown by

- a. Stoney
- c. Millikan

- b. J.J Thomson
- d. Rutherford

THE STAR INSTITUTE

